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Topic 2

The given statement is true. Percentage of Carbon contained in a fuel is needed to calculate the CO_2 emissions.

We know that

$$\underset{12}{C} + \underset{32}{O_2} \rightarrow \underset{44}{CO_2}$$

i.e. 12Kg of Carbon with complete combustion forms 44Kg of Carbon Dioxide.

$\Rightarrow$ 1Kg of Carbon with complete combustion forms 44/12 = 3.67 Kg of Carbon Dioxide.

CO_2 Liberated from Gasoline

Indian Gasoline contains 85.4% of Carbon.

Hence 1 Kg of Indian Gasoline contains 0.854 Kg of Carbon.

With complete combustion, 1 Kg of Indian Gasoline forms 0.854 * 3.67 = 3.134 Kg of Carbon Dioxide. …… (**A)**

CO_2 Liberated from LPG

Indian LPG contains 82.3% of Carbon.

Hence 1 Kg of Indian LPG contains 0.823 Kg of Carbon.

With complete combustion, 1 Kg of Indian LPG forms 0.823 * 3.67 = 3.020 Kg of Carbon Dioxide. …… (**B**)

Comparing (**A**) and (**B**), it can be concluded that **Carbon Dioxide liberated from complete combustion of 1 Kg of Indian Gasoline is more than that from 1 Kg of Indian LPG.**