

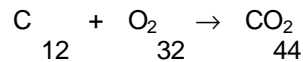


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Topic 2

The given statement is true. Percentage of Carbon contained in a fuel is needed to calculate the CO₂ emissions.

We know that



i.e. 12Kg of Carbon with complete combustion forms 44Kg of Carbon Dioxide.

⇒ 1Kg of Carbon with complete combustion forms $44/12 = 3.67$ Kg of Carbon Dioxide.

CO₂ Liberated from Gasoline

Indian Gasoline contains 85.4% of Carbon.

Hence 1 Kg of Indian Gasoline contains 0.854 Kg of Carbon.

With complete combustion, 1 Kg of Indian Gasoline forms $0.854 * 3.67 = 3.134$ Kg of Carbon Dioxide. (A)

CO₂ Liberated from LPG

Indian LPG contains 82.3% of Carbon.

Hence 1 Kg of Indian LPG contains 0.823 Kg of Carbon.

With complete combustion, 1 Kg of Indian LPG forms $0.823 * 3.67 = 3.020$ Kg of Carbon Dioxide. (B)

Comparing (A) and (B), it can be concluded that **Carbon Dioxide liberated from complete combustion of 1 Kg of Indian Gasoline is more than that from 1 Kg of Indian LPG.**